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## Structural characterization of some fluoroquinolone drugs

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### Abstract

The vibrational spectroscopy, such as FTIR has been used to measure the vibrational modes of fluoroquinolones, provides information about structural differences of its individual members. From the interpreted spectral data Ofloxacin and Ciprofloxacin have been distinguished by the presence of different substituents in their parent nucleus. FTIR study provides the most direct and definitive identification of fluoroquinolones and offer a means for qualitative analysis of pure drugs.

**Keywords:** Structural characterization, fluoroquinolone drugs

### Introduction

Fluoroquinolones have been associated with a significant number of serious adverse drug reactions such as tendon damage <sup>[1]</sup> and peripheral neuropathy, such as reactions may manifest long after therapy had been completed and in severe cases may result in lifelong disabilities <sup>[2]</sup>. They are associated with severe psychiatric <sup>[3, 4]</sup> adverse reactions. The reaction was detailed within Stephen Fried's book Bitter Pills (1999). Hepatotoxicity has also been reported with the use of some fluoroquinolones <sup>[5, 6]</sup>.

Ofloxacin, one of the fluoroquinolones, has not been shown to be effective in the treatment of syphilis <sup>[7, 8]</sup> like other fluoroquinolones. Some fluoroquinolones are effective for the treatment of certain sexually transmitted diseases due to bacterial resistance. Some of them are broad spectrum antibiotic that is active against both gram-positive and Gram-negative bacteria. They act by inhibiting DNA gyrase a type-II topoisomerase and topoisomerase-IV <sup>[9]</sup> which is an enzyme necessary to separate replicated DNA thereby inhibiting bacterial cell division. The fluoroquinolones interfere with DNA replication by inhibiting an enzyme complex called DNA gyrase. This can also affect mammalian cell replication. Some fluoroquinolones may cause injury to the chromosome of eukaryotic cells <sup>[10]</sup>.

### Materials and Methods

#### Materials

Pure samples of ofloxacin (C<sub>18</sub>H<sub>20</sub>FN<sub>3</sub>O<sub>4</sub>) and ciprofloxacin (C<sub>17</sub>H<sub>18</sub>FN<sub>3</sub>O<sub>3</sub>) were obtained from external agency. The drugs were of 99.8% to 98.0% purity.

#### Methods

FTIR Spectroscopy is an important analytical technique which detects various characteristic functional groups in molecules of any matter. On interaction of an infrared light with the matter, chemical bonds would stretch, contract and bend, and as a result each chemical functional group tends to absorb infrared radiation in a specific wavelength range regardless of the structure of the rest of the molecule. Based on this principle, functional groups present in composite materials are identified. It is performed in a FTIR

Spectrophotometer interfaced with infrared (IR) microscope operated in reflectance mode. The microscope is equipped with a video camera, a liquid Nitrogen-cooled Mercury Cadmium Telluride (MCT) detector and a computer controlled translation stage, programmable in the x and y directions. The FTIR imaging in the present investigation was carried out using a Perkin Elmer Spectrum RX. Here KBr pellet method was used for sample preparation for FTIR study<sup>13</sup>. The spectra were collected in the 400 cm<sup>-1</sup> to 4000 cm<sup>-1</sup> region with 8 cm<sup>-1</sup> resolution, 60 scans and beam spot size of 10 μm-100 μm.

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## Results and Discussion

The infrared spectra are recorded on Fourier Transform Spectrometer in the mid-infrared region (MIR) within the range (400-4500  $\text{cm}^{-1}$ ) [11]. Due to the complex interaction of atoms within the molecule, IR absorption of the functional groups may vary over a wide range. However, it has been found that many functional groups give characteristic IR absorption at specific narrow frequency range. Multiple functional groups may absorb at one particular frequency range but a functional group often gives rise to several characteristic absorptions. Thus, the spectral interpretations should not be confined to one or two bands only; actually, the whole spectrum should be examined.

While the FTIR band at 4000-1300  $\text{cm}^{-1}$  represented functional group region, the appearance of strong absorption bands in the region of 4000 to 2500  $\text{cm}^{-1}$  was due to stretching vibrations between hydrogen and some other atoms with a mass of 19 or less.

The O-H and N-H stretching frequencies were in the 3700 to 2500  $\text{cm}^{-1}$  region with various intensities. Hydrogen bonding has a significant influence on the peak shape and intensities, generally causing peak broadening and shifts in absorption to lower frequencies. The C-H stretching

vibration occurred [12] in the region of 3300 to 2800  $\text{cm}^{-1}$ .

In FTIR spectra of Oflox, one prominent characteristic peak was found between 3050 and 3000  $\text{cm}^{-1}$ , which was assigned to stretching vibration of OH group and intramolecular hydrogen bonding. This band also suggested the NH stretching vibration of the imino-moiety of piperazinyl groups which was less prominent due to intense OH stretching vibration. The peak at 2700  $\text{cm}^{-1}$  was assigned to  $\nu\text{CH}_3$  of methyl group. The band at 1750-1700  $\text{cm}^{-1}$  represented the acidic carbonyl C=O stretching i.e.,  $\nu\text{C}=\text{O}$  [13]. The peak at 1650 to 1600  $\text{cm}^{-1}$  was assigned to  $\nu\text{N-H}$  bending vibration of quinolones. The 1550 to 1500  $\text{cm}^{-1}$  represented the  $\nu\text{CH}_2$  of the aromatic ring. The band at 1450-1400  $\text{cm}^{-1}$  was assigned to the stretching vibration of  $\text{CH}_2$  confirming the presence of methylene group in benzoxazine ring. The peak at 1400-1350  $\text{cm}^{-1}$  represented the bending vibration of hydroxyl group. The band at the 1250 to 1200  $\text{cm}^{-1}$  suggested the stretching vibration of oxo group. In addition, a strong absorption peak between 1050 and 1000  $\text{cm}^{-1}$  was assigned to C-F group. The band at 900-800  $\text{cm}^{-1}$  represented the out of plane bending vibration of double bonded enes or =CH groups [14, 15].

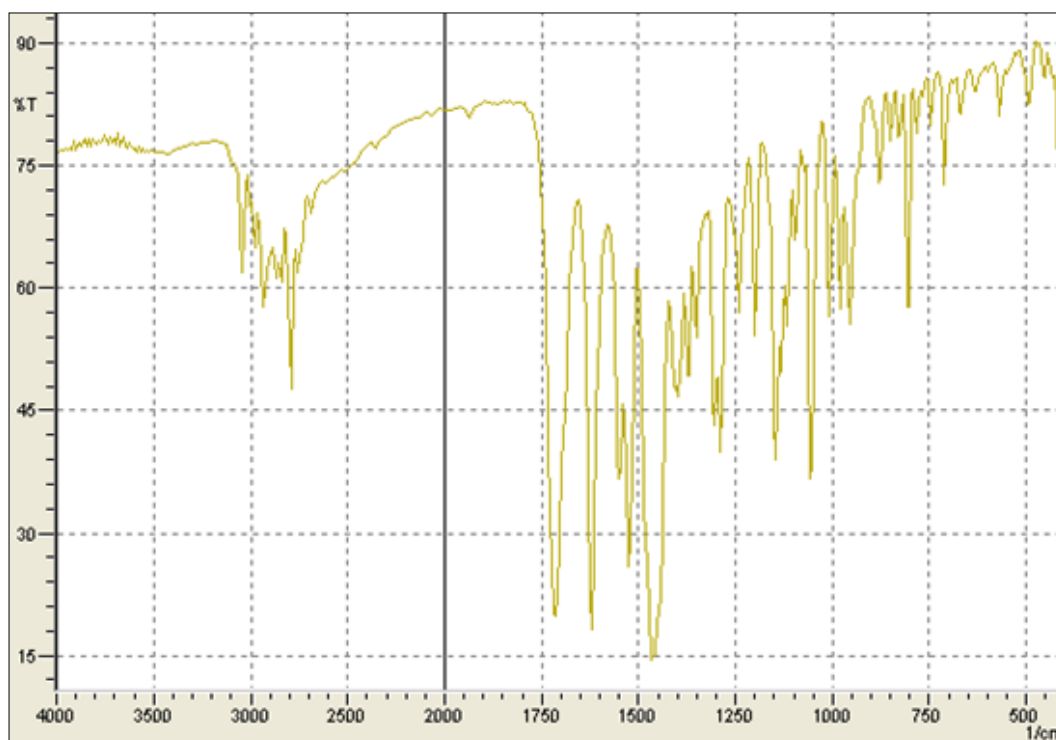


Fig 1: IR spectrum of Ofloxacin

In FTIR spectra of Cipro, one prominent characteristic peak was found between 3500-3450  $\text{cm}^{-1}$ , which was assigned to stretching vibration of OH group and intermolecular hydrogen bonding. The possible peak of imino moiety of the piperazinyl group was less prominent due to intense OH stretching vibration. Another band at 3000-2950  $\text{cm}^{-1}$  represented the alkene and aromatic C-H stretching, mainly  $\nu=\text{CH}$ .

The peak at 2900  $\text{cm}^{-1}$  was assigned to C-H stretching vibration of cyclopropyl group. The 1950 to 1450  $\text{cm}^{-1}$

region exhibited FTIR absorption from a wide variety of double-bonded functional groups. The band at 1750 to 1700  $\text{cm}^{-1}$  represented the carbonyl C=O stretching i.e.,  $\nu\text{C}=\text{O}$ . The peak at 1650 to 1600  $\text{cm}^{-1}$  was assigned to quinolones. The band at the 1450 to 1400  $\text{cm}^{-1}$  represented  $\nu\text{CO}$  and at 1300 to 1250  $\text{cm}^{-1}$  suggested bending vibration of O-H group which proved the presence of carboxylic acid. A strong absorption peak between 1050 and 1000  $\text{cm}^{-1}$  was assigned to C-F group [16, 17].

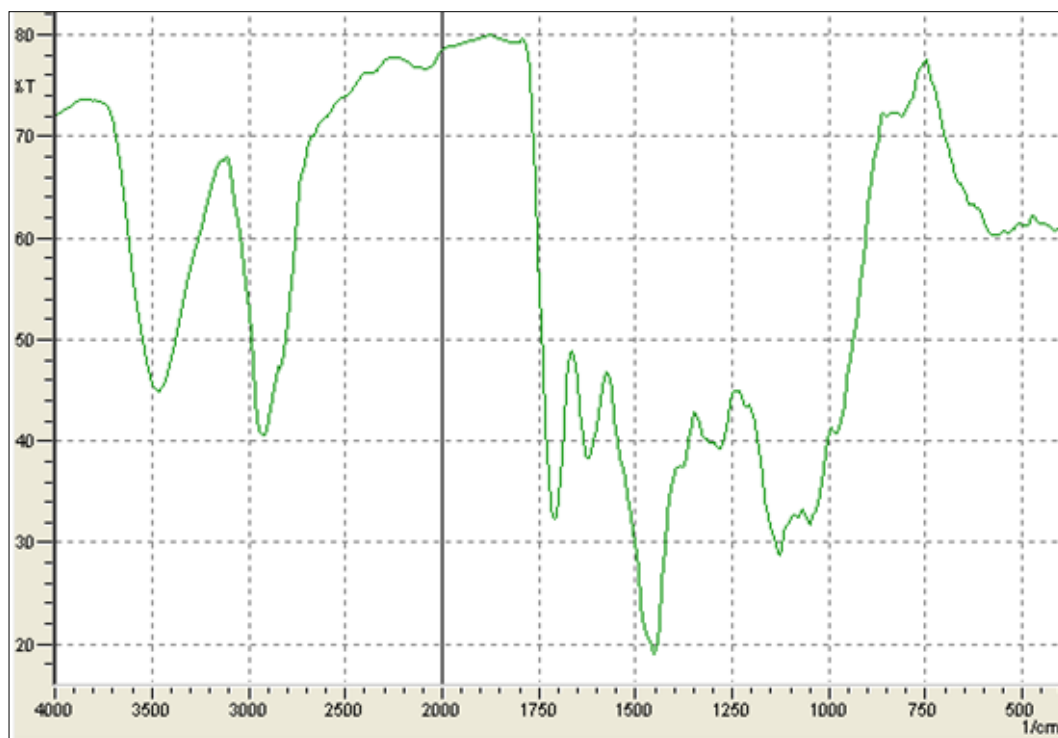
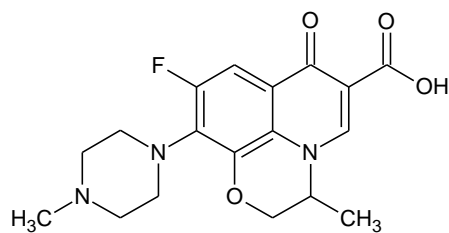


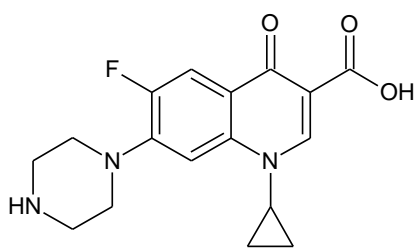
Fig 2: IR spectrum of Ciprofloxacin

### Conclusion

On the basis of FTIR spectral analysis of Ofloxacin and ciprofloxacin, the structures of both fluoroquinolones may be written as:



Ofloxacin



Ciprofloxacin

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